

Class XII Session 2024-25
Subject - Chemistry
Sample Question Paper - 10

Time Allowed: 3 hours

Maximum Marks: 70

General Instructions:

Read the following instructions carefully.

1. There are **33** questions in this question paper with internal choice.
2. SECTION A consists of 16 multiple-choice questions carrying 1 mark each.
3. SECTION B consists of 5 very short answer questions carrying 2 marks each.
4. SECTION C consists of 7 short answer questions carrying 3 marks each.
5. SECTION D consists of 2 case-based questions carrying 4 marks each.
6. SECTION E consists of 3 long answer questions carrying 5 marks each.
7. **All questions are compulsory.**
8. **Use of log tables and calculators is not allowed.**

Section A

- Chlorobenzene is formed by the reaction of chlorine with benzene in the presence of AlCl_3 . Which of the following species attacks the benzene ring in this reaction?
a) AlCl_3 b) $[\text{AlCl}_4]^-$
c) Cl^+ d) Cl^-
- Which of the following reactions of glucose can be explained only by its cyclic structure?
a) Glucose is oxidised by nitric acid to gluconic acid.
b) Pentaacetate of glucose does not react with hydroxylamine.
c) Glucose reacts with hydroxylamine to form an oxime.
d) Glucose forms pentaacetate.
- CH_3CONH_2 on reaction with NaOH and Br_2 in alcoholic medium gives:
a) CH_3NH_2 b) $\text{CH}_3\text{CH}_2\text{NH}_2$
c) CH_3COONa d) $\text{CH}_3\text{CH}_2\text{Br}$
- The compound which forms acetaldehyde when heated with dilute NaOH is:
a) 1, 2 dichloroethane b) 1, 1, 1 trichloroethane
c) 1 chloroethane d) 1, 1 dichloroethane
- The reaction $A \rightarrow B$ is a second order process when the initial concentration of A is 0.50 M, the half life is 8.0 minutes. What is the half life if the initial concentration of A is 0.10 M?

6. Match the items given in column I with that in column II: [1]

a) (a) - (i), (b) - (ii), (c) - (iii), (d) - (iv) b) (a) - (iv), (b) - (ii), (c) - (iii), (d) - (i)

c) (a) - (i), (b) - (ii), (c) - (iii), (d) - (iv) d) (a) - (ii), (b) - (iv), (c) - (i), (d) - (iii)

- C_2H_5
- a) 1-Bromo-2-ethyl-2-methylethane b) 1-Bromo-2-methylbutane
- c) 2-Methyl-1-bromobutane d) 1-Bromo-2-ethylpropane

- a) H₂S b) O₂
c) Cl₂ d) N₂

- $$A_{(g)} + B_{(g)} \rightarrow C_{(g)} + D_{(g)}$$

What are the orders with respect to $A_{(g)}$ and $B_{(g)}$?

- | | |
|---|---|
| a) Order with respect to $A_{(g)}$ - Second | b) Order with respect to $A_{(g)}$ - Zero |
| Order with respect to $B_{(g)}$ - First | Order with respect to $B_{(g)}$ - Second |
| c) Order with respect to $A_{(g)}$ - First | d) Order with respect to $A_{(g)}$ - Second |
| Order with respect to $B_{(g)}$ - Zero | Order with respect to $B_{(g)}$ - Zero |

- a) 1-aminoethane
b) Ethanamide
c) 1-aminopropane
d) Ethanoic acid

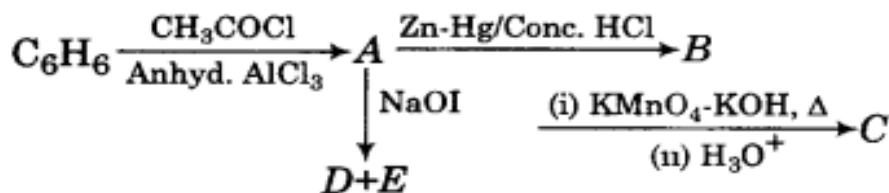
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OR

What happens when

- a pressure greater than osmotic pressure is applied on the solution side separated from solvent by a semipermeable membrane?
- acetone is added to pure ethanol?

21. Write the structures of A, B, C, D and E in the following reactions. [2]



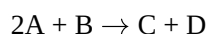
Section C

22. Zinc rod is dipped in 0.1 M solution of ZnSO_4 [3]

The salt is 95% dissociated at its dilution at 298 K. Calculate the electrode potential. Given:

$$E^0(\text{Zn}^{2+}/\text{Zn}) = -0.76$$

23. The following results have been obtained during the kinetic studies of the reaction. [3]



Experiment	[A]/mol L ⁻¹	[B]/mol L ⁻¹	Initial rate of formation of D/mol L ⁻¹ min ⁻¹
I	0.1	0.1	6.0×10^{-3}
II	0.3	0.2	7.2×10^{-2}
III	0.3	0.4	2.88×10^{-1}
IV	0.4	0.1	2.40×10^{-2}

Determine the rate law and the rate constant for the reaction.

24. Give structures of the products you would expect when each of the following alcohols [3]

- Butan-1-ol
- 2-Methylbutan-2-ol react with
 - $\text{HCl} - \text{ZnCl}_2$
 - HBr and
 - SOCl_2

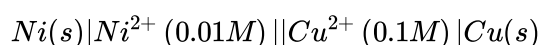
OR

Explain how does the -OH group attached to a carbon of benzene ring activate it towards electrophilic substitution?

25. Draw structures of the following derivatives: [3]

- 2,4-Dinitrophenylhydrazone of benzaldehyde
- Cyclopropanone oxime
- Acetaldehydedimethylacetal

26. Calculate the e.m.f at 25°C for the following cell: [3]



$$\text{Given: } E^0_{\text{Ni}^{2+}/\text{Ni}} = -0.25 \text{ V}$$

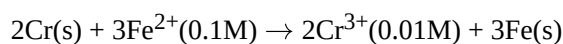
$$E_{Cu^{2+}/Cu}^0 = +0.34V$$

[1F = 96,500 Cmol⁻¹]. Calculate the maximum work that can be accomplished by the operation of this cell

27. a. Out of C1CCCCC1Cl and C1CCCCC1CCl, which one is more reactive towards S_N2 reaction and why? [3]
- b. Out of c1ccccc1Cl and O=[N+]([O-])c1ccc(Cl)cc1, which one is more reactive towards nucleophilic substitution reaction and why?

- c. Out of CC(C)CCO and CC(C)(O)CC, which one is optically active and why?

28. Calculate the emf of the following cell at 298 K. [3]

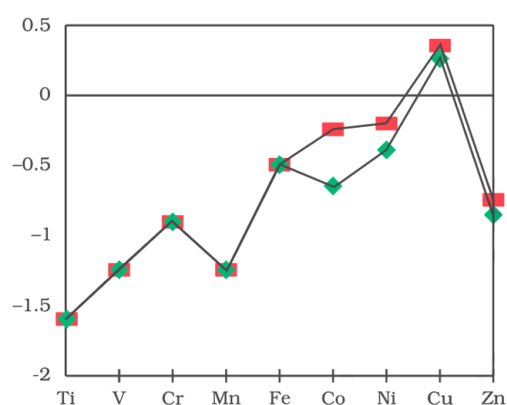


Given, $E_{Cr^{3+}/Cr}^0 = -0.74V$, $E_{Fe^{2+}/Fe}^0 = -0.44V$

Section D

29. Read the following text carefully and answer the questions that follow: [4]

The unique behaviour of Cu, having a positive E^0 accounts for its inability to liberate H₂ from acids. Only oxidising acids (nitric and hot concentrated sulphuric) react with Cu, the acids being reduced. The stability of the half-filled d sub-shell in Mn²⁺ and the completely filled d¹⁰ configuration in Zn²⁺ are related to their E^0 values, whereas E^0 for Ni is related to the highest negative $\Delta_{hyd}H^0$. An examination of the $E_{(M^{3+}/M^{2+})}^0$ values the low value for Sc reflects the stability of Sc³⁺ which has a noble gas configuration. The comparatively high value for Mn shows that Mn²⁺(d⁵) is particularly stable, whereas a comparatively low value for Fe shows the extra stability of Fe³⁺(d⁵). The comparatively low value for V is related to the stability of V²⁺ (half-filled t_{2g} level).



- Why Zn has high value for M³⁺/M²⁺ Standard Electrode Potentials? (1)
- Transition metals, despite high E^0 oxidation, are poor reducing agents. Justify. (1)
- Why is Cr²⁺ reducing and Mn³⁺ oxidising when both Cr and Mn have d⁴ configuration? (2)

OR

Why Cu²⁺ is more stable than Cu⁺? (2)

30. Read the following text carefully and answer the questions that follow: [4]

Vapour pressure of a liquid or a solution is the pressure exerted by the vapour in equilibrium with the liquid or solution at a particular temperature. It depends upon the nature of the liquid and temperature. The non-volatile solute in solution reduces the escaping tendency of the solvent molecules in the vapour phase because some of

the solute particles occupy the positions of the solvent molecules on the liquid surface. The relative lowering of the vapour pressure of a solution containing a non-volatile solute is equal to the mole fraction of the solute in the solution. This is also known as Raoult's law. However, for solutions of volatile solutes, the vapour pressure of a component in a solution at a given temperature is equal to the mole fraction of that component in the solution multiplied by the vapour pressure of that pure component. The solutions in which each component obeys Raoult's law is called an ideal solution. For ideal solutions ΔH_{mixing} and ΔV_{mixing} are also zero. Practically no solution is ideal. A non-ideal solution is that solution in which solute and solvent molecules interact with one another with a different force than the forces of interaction between the molecules of the pure components. There are two types of non-ideal solutions, showing positive deviations and negative deviations from ideal behaviour. If for the two components A and B, the forces of interaction between A and B molecules are less than the A-A and B-B interactions, the non-ideal solutions have positive deviations. On the other hand, if the forces of interaction between A and B molecules are more than the A-A and B-B interactions, the non-ideal solutions have negative deviations.

- What is the mole fraction of A in solution obeying result's law if the vapour pressure of a pure liquid A is 40 mm of Hg at 300 K. The vapour pressure of this liquid in solution with liquid B is 32 mm of Hg? (1)
- Vapour pressure of a solution of heptane & octane is given by the equation: (1)
 $P(\text{sol.})(\text{mm Hg}) = 35 + 65x$, where x is the mole fraction of heptane. Calculate the vapour pressure of pure octane. (2)
- What is the value of ΔV_{mixing} and ΔH_{mixing} for non-ideal solution showing negative deviation?

OR

Acetic acid + pyridine, the mixture is an example of which type of solution? (2)

Section E

- Attempt any five of the following: [5]
 - Which sugar is present in milk? [1]
 - Why must vitamin C be supplied regularly in diet? [1]
 - What type of substance is phenylalanine hydroxylate? Write its importance. [1]
 - Define denaturation of protein. What is the effect of denaturation on the structure of protein? [1]
 - Name the linkage connecting monosaccharide units in polysaccharides. [1]
 - How do you explain the presence of all six carbon atoms in glucose in a straight chain? [1]
 - Write the reactions showing the presence of following in the open structure of glucose: [1]
 - a carbonyl group
 - chain with six carbon atoms
- Amongst the following, the most stable complex is: [5]
 - $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$
 - $[\text{Fe}(\text{NH}_3)_6]^{3+}$
 - $[\text{Fe}(\text{C}_2\text{O}_4)_3]^{3-}$
 - $[\text{FeCl}_6]^{3+}$

b. What will be the correct order for the wavelength of absorption in the visible region for the following:
 $[\text{Ni}(\text{NO}_2)_6]^{4-}$, $[\text{Ni}(\text{NH}_3)_6]^{2+}$, $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$

OR



FeSO_4 solution mixed with $(\text{NH}_4)_2\text{SO}_4$ solution in 1:1 molar ratio gives the test of Fe^{2+} ion but CuSO_4 solution mixed with aqueous ammonia in 1:4 molar ratio does not give the test of Cu^{2+} ion. Explain why?

33. State the reactions and reaction conditions for the following conversions :

[5]

- i. Benzene diazonium chloride to nitrobenzene.
- ii. Aniline to benzene diazonium chloride.
- iii. Ethylamine to methylamine.

OR

- i. Tert-Butylamine cannot be prepared by the action of NH_3 on tert-butyl bromide. Explain why?
- ii. Suggest a convenient method for the preparation of tert-butylamine.



Solution

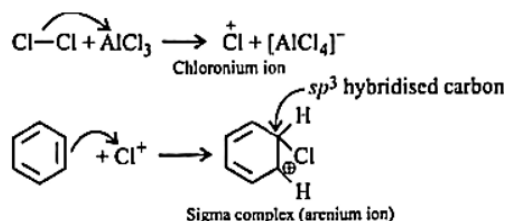
Section A

1.

(c) Cl^+

Explanation:

Aluminum chloride (AlCl_3) is a Lewis acid catalyst and works in the same way as FeCl_3 does. Benzene (C_6H_6) is converted into chlorobenzene by chlorination of benzene in the presence of AlCl_3 . The reaction occurs by an electrophilic substitution reaction. Cl_2 forms a coordination complex with AlCl_3 , forming $\text{Cl}^+\text{AlCl}_4^-$ complex, which gives a slight positive charge to Cl , and AlCl_4^- is negatively charged. This Cl^+ then reacts with the aromatic double bond of the benzene ring to form an additional product, followed by deprotonation to form chlorobenzene and AlCl_3 and HCl as the side products.



2.

(b) Pentaacetate of glucose does not react with hydroxylamine.

Explanation: The pentaacetate of glucose does not react with the hydroxylamine indicating the absence of free $-\text{CHO}$ group. This property of the glucose can be explained only by its own cyclic structure.

3.

(a) CH_3NH_2

Explanation: Product formed is CH_3NH_2

4.

(d) 1, 1 dichloroethane

Explanation: $\text{CH}_3\text{CHCl}_2 + \text{OH}^- \rightarrow \text{CH}_3\text{CH}(\text{OH})_2 \rightarrow \text{CH}_3\text{CHO} + \text{H}_2\text{O}$

Gem diols like $(\text{CH}_3\text{CH}(\text{OH})_2)$ are generally not stable. The 2 $-\text{OH}$ group attached to the same C removes H_2O and forms carbonyl compounds.

5.

(a) 40.0 minutes

Explanation: For second-order reaction:

$$t_{1/2} = \frac{2k}{[R]} \Rightarrow k = \frac{t_{1/2}[R]}{2}$$

Applying this equation,

$$\frac{t_{1/2}[R]}{2} = \frac{t'_{1/2}[R']}{2}$$

$$\frac{8.0 \times 0.50}{2} = \frac{t'_{1/2} \times 0.10}{2}$$

$$t'_{1/2} = \frac{8 \times 0.50}{0.10} = 40$$

6.

(d) (a) - (ii), (b) - (iv), (c) - (i), (d) - (iii)

Explanation: (a) - (ii), (b) - (iv), (c) - (i), (d) - (iii)

7.

(b) 1-Bromo-2-methylbutane

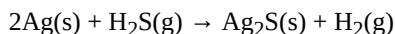
Explanation: First, we need to identify the longest carbon chain. Once we do that, the actual structure should read $\text{CH}_3-\text{CH}_2-\text{CH}(\text{CH}_3)-\text{CH}_2-\text{Br}$. $-\text{Br}$, the functional halide group is attached to the first carbon atom (1- Bromo), so we start the



numbering from that position. The methyl group branch is bond to the second carbon atom in the chain(2-methyl). The number of carbons in the unbranched parent chain is four, thus giving the name butane. The IUPAC is named 1-Bromo-2-methylbutane.

8. (a) H_2S

Explanation: Silver ornaments turns black coming in contact with H_2S due to formation of Ag_2S . The chemical equation for this change can be represented as given below:



9.

(d) Order with respect to $\text{A}_{(\text{g})}$ - Second

Order with respect to $\text{B}_{(\text{g})}$ - Zero

Explanation: Order with respect to $\text{A}_{(\text{g})}$ - Second

Order with respect to $\text{B}_{(\text{g})}$ - Zero

10. (a) 1-aminoethane

Explanation: 1-aminoethane

11.

(b) ethers

Explanation: The Williamson ether synthesis is an organic reaction, forming an ether from an organohalide and deprotonated alcohol (alkoxide). This reaction was developed by Alexander Williamson in 1850. Typically it involves the reaction of an alkoxide ion with a primary alkyl halide via an $\text{S}_{\text{N}}2$ reaction.



12.

(b) Aluminium chloride reacts with Aniline

Explanation: AlCl_3 being a lewis acid reacts with the lone pair of $-\text{NH}_2$ group of aniline forming an adduct ($\text{C}_6\text{H}_5\text{NH}_2^+ \text{AlCl}_3^-$) which deactivates the benzene system hence no friedal craft reaction occurs.

13.

(b) Both A and R are true but R is not the correct explanation of A.

Explanation: Cellulose is a linear condensation polymer of β -glucose. Because of linear structure, cellulose can be easily converted into fibres.

14. (a) Both A and R are true and R is the correct explanation of A.

Explanation: The release of hydrogen from α -carbon of aldehydes or ketones results in the formation of carbanion which is resonance stabilised. It is called enolate. Enolate ion can act as as a nucleophile and can attack the carbonyl group of the other molecule, i.e. giving condensation reactions.

15.

(c) Assertion (A) is true, but Reason (R) is false.

Explanation: The amine formed reacts further with more of alkyl halide to give 2° and 3° amines and finally quaternary salt

16.

(b) Both A and R are true but R is not the correct explanation of A.

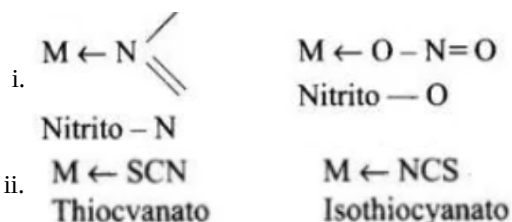
Explanation: Removal of two molecules of water gives a product which tautomerises to yield acrolein -an α, β -unsaturated aldehyde.

Section B

17. Linkage isomerism- arises in a coordinated compound containing ambidentate ligand

Examples:





18. The high melting points of transition metals are due to strong metallic bonds between the atoms of these elements. This is because of involvements of a greater number of electrons in the interatomic metallic bonding from $(n-1)d$ orbitals in addition to ns electrons in the interatomic molecular bonding.

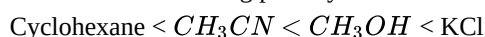
19. Answer the following:

(i) First order.

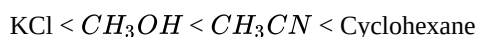
(ii) Elementary reactions have the same value of order and molecularity because the elementary reaction proceeds in a single step.

20. n-octane is a non-polar solvent. Therefore, the solubility of a non-polar solute is more than that of a polar solute in the n-octane.

The order of increasing polarity is:



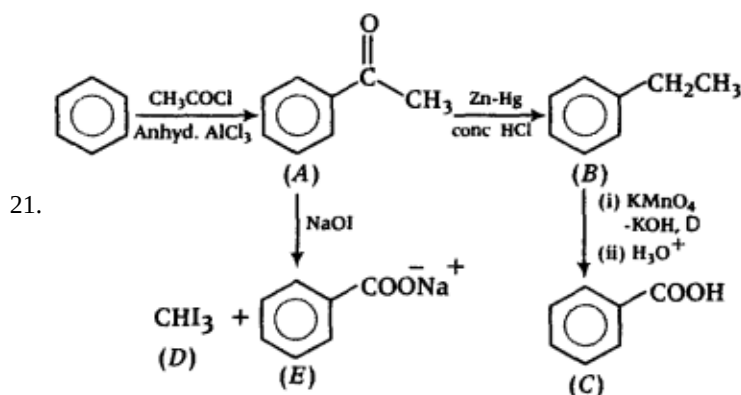
Therefore, the order of increasing solubility is:



OR

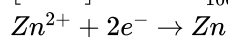
i. Reverse osmosis (direction of osmosis is reversed) occurs i.e. net flow of the solvent is from solution to solvent. Hence greater pressure is applied to the solvent molecules from the solution side would move into the pure solvent through a semipermeable membrane.

ii. When acetone is added to the ethanol, molecules of acetone get in between the molecules of ethanol and break some of the hydrogen bond, which weakens the intermolecular attractive forces resulting in an increase in vapor pressure of the mixture than ethanol in a pure state.



Section C

22. $[Zn^{2+}] = 0.1 \times \frac{95}{100} = 0.095M$



$E_{(Zn^{2+}/Zn)} = E_{(Zn^{2+}/Zn)}^0 - \frac{0.0591}{2} \log \frac{1}{[Zn^{2+}]}$

$= -0.76V - \frac{0.0591}{2} \log \frac{1}{0.095}$

$= -0.76V - \frac{0.0591}{2} [\log 1000 - \log 95]$

$= -0.76 - \frac{0.0591}{2} [3.000 - 1.9777]$

$= -0.76V - \frac{0.0591}{2} \times 1.0223$

$= -0.76V - \frac{0.0604}{2} = 0.76 - 0.0302$

$= -0.7902V$

23. From experiments I and IV, it may be noted that [B] is same but [A] has been made four times, the rate of reaction has also becomes four times. This means w.r.t A

Rate $\propto [A]$

From experiments II and III it may be noted that [A] is kept same and [B] has been doubled, the rate of reaction has become four times. This means w.r.t B,

Rate $\propto [B]^2$

Combining (i) and (ii), we get the rate law for the given reaction as:

$$\text{Rate} = k[A][B]^2$$

Thus order w.r.t A = 1

Order w.r.t B = 2

The overall order of the reaction

$$= 1 + 2 = 3$$

The rate constant and its units can be calculated from the data of each experiment using the expression.

$$k = \frac{\text{rate}}{[A][B]^2}$$

$$= \frac{\text{mol L}^{-1} \text{min}^{-1}}{(\text{mol L}^{-1})(\text{mol L}^{-1})^2}$$

$$= \text{mol}^{-2} \text{L}^{-2} \text{min}^{-1}$$

$$\text{Expt. } k = \text{mol}^{-2} \text{L}^{-2} \text{min}^{-1}$$

$$\text{i. } \frac{6.0 \times 10^{-3}}{0.1 \times (0.1)^2} = 6.0$$

$$\text{ii. } \frac{7.2 \times 10^{-2}}{0.3 \times (0.2)^2} = 6.0$$

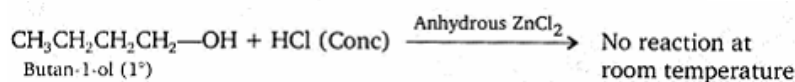
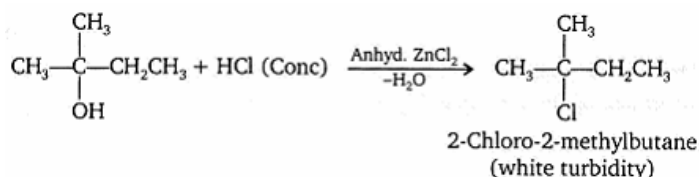
$$\text{iii. } \frac{2.88 \times 10^{-1}}{0.3 \times (0.4)^2} = 6.0$$

$$\text{iv. } \frac{2.4 \times 10^{-2}}{0.4 \times (0.1)^2} = 6.0$$

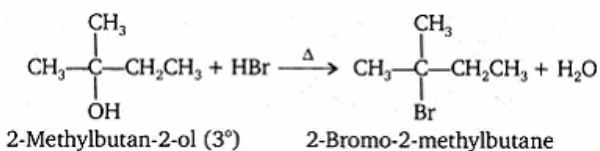
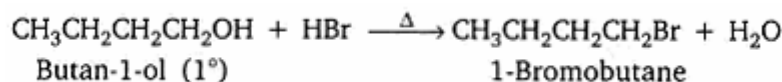
$$\text{Rate constant } k = 6.0 \text{ mol}^{-2} \text{L}^{-2} \text{min}^{-1}$$

24. i. a. With HCl - ZnCl₂ (lucas reagent) 2-Methylbutane-2-ol

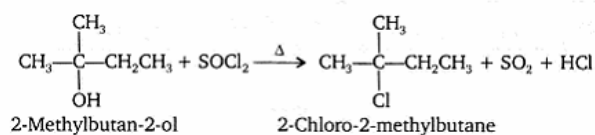
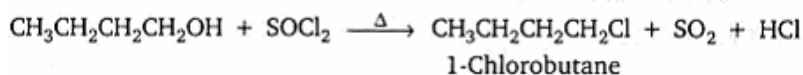
ii. Being a 3° alcohol, reacts with Lucas reagent to produce turbidity immediately due to the formation of insoluble tert-alkyl chloride while butane-1-ol (i) being a 1° alcohol does not react with Lucas reagent at room temperature.



a. Both the alcohols react to produce the corresponding alkyl bromides.

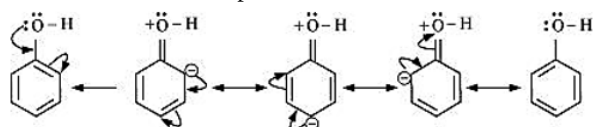


b. Both the alcohols react to form the corresponding alkyl chlorides.

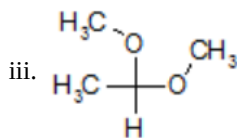
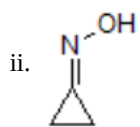
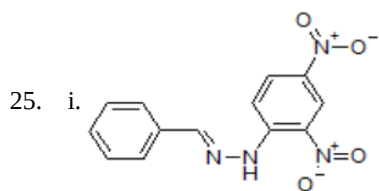


OR

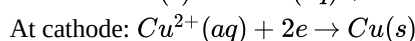
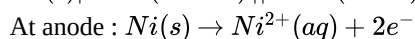
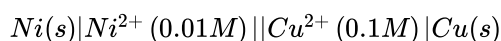
The -OH group is an electron-donating group. Thus, it increases the electron density in the benzene ring as shown in the given resonance structure of phenol.



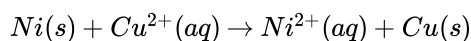
As a result, the benzene ring is activated towards electrophilic substitution.



26. Cell reaction is



Net cell reaction



$$E_{cell} = E_{cell}^0 - \frac{0.0591}{2} \log \frac{[Ni^{2+}]}{[Cu^{2+}]}$$

$$= [+0.34V - (-0.25V)] - \frac{0.0591}{2} \log \frac{1}{10}$$

$$= 0.59V - \frac{0.0591}{2} \times -1$$

$$= 0.59 + 0.0295$$

$$= 0.6195V$$

$$\Delta G = -nEF$$

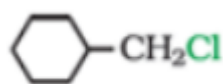
$$= -2 \times 0.6195V \times 96500C/mol$$

$$= -119563.5J/mol$$

$$\Delta G = -119.5635KJ/mol$$

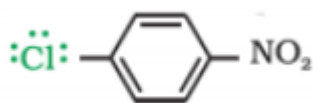
$$-\Delta G = W_{max} = 119.5635KJ/mol$$

27. a.

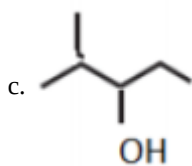


is more reactive towards S_N2 because it is a primary halide.

b.

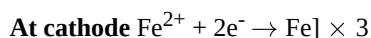
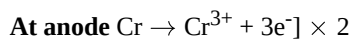


is more reactive towards nucleophilic substitution reaction because it is more reactive due to the presence of electron withdrawing - NO_2 group.

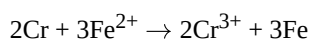


It is optically active due to chiral carbon.

28. Since oxidation of Cr is taking place in the given reaction, the chromium electrode is anode and as Fe is reduced in the reaction, Fe electrode is the cathode. The half-cell reactions are as follows.



Overall reaction



$$E_o = E_{\text{cathode}} - E_{\text{anode}} = -0.44 - (-0.74) = 0.3V$$

$$E = E^o - \frac{0.0591}{n} \log \frac{[Cr^{3+}]^2}{[Fe^{2+}]^3}$$

Here, n = number of electrons transferred, i.e. equal to 6.

$$= 0.3 - \frac{0.0591}{6} \log \frac{[0.01]^2}{[0.1]^3}$$

$$= 0.309 \approx 0.31$$

Section D

29. i. Due to the removal of an electron from the stable d^{10} configuration of Zn^{2+} .
 ii. Transition metals despite having high E° oxidation, are poor reducing agents because of their high heat of vaporization, high ionisation energies and low heats of hydration.
 iii. Cr^{2+} is reducing as its configuration changes from d_4 to d_3 , the having a half-filled t_{2g} level. On the other hand, the change from Mn^{3+} to Mn^{2+} results in the half-filled (d_5) configuration which has extra stability.

OR

The Stability of Cu^{2+} is more than Cu^+ as stability depends on the hydration energy of the ions when they bond to the water molecules. The Cu^{2+} ion has a greater charge density than Cu^+ ion and thus forms much stronger bonds releasing more energy.

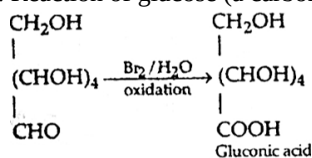
30. i. $P_A = x_A \times P_A^\circ$
 $32 = x_A \times 40$
 $x_A = \frac{32}{40}$
 $x_A = 0.8$
 ii. For pure octane, $x = 0$
 $\therefore p(\text{sol.})(\text{mm Hg}) = P(\text{octane}) = 35 + 65 \times 0 = 35 \text{ mm of Hg}$
 iii. The value of ΔV_{mixing} and ΔH_{mixing} is negative.
 iv. It is an example of Non-ideal solution.

Section E

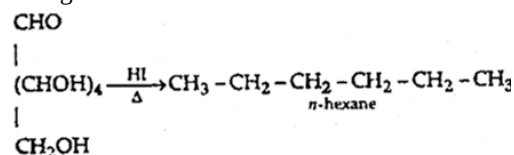
31. Attempt any five of the following:
 (i) Lactose is the type of sugar that occurs naturally in milk. It is found in the milk of animals such as cows and goats, as well as human breast milk.
 (ii) Vitamin 'C' is water soluble vitamin and hence excess of it is readily excreted in the urine so, it cannot be stored in our body and hence, it should be regularly supplied in the diet.
 (iii) It is an enzyme whose deficiency causes mental retardation.
 (iv) When a protein is subjected to a change in temperature or chemical change then it loses its biological activity.
 2° and 3° structures are destroyed but 1° structure remains intact.
 (v) The monosaccharide units are linked through the glycosidic linkage in the polysaccharide.
 (vi) Glucose on prolonged heating with HI and red phosphorus gives n-hexane HI (excess) suggest that all six carbon atom are in straight chain.



- (vii) i. Reaction of glucose (a carbonyl group)



- ii. Straight chain with six carbon atoms



32. a. Complexes containing didentate or polydentate ligands are more stable than those containing monodentate ligands. In each of the given complex, Fe is in +3 state.

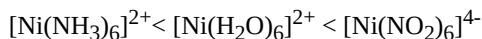


As $C_2O_4^{2-}$ is didentate chelating ligand, hence is the most stable complex.

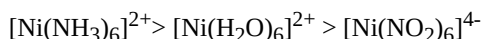
b. As metal ion is fixed, the wavelength of absorption will be decided by the field strengths (CFSE values) of the ligands.

From the spectrochemical series, the order of CFSE is: $H_2O < NH_3 < NO_2$

Thus, the energies absorbed for excitation will be in the order:



As wavelength and energy are inversely related. The wavelengths absorbed will be in the opposite order:

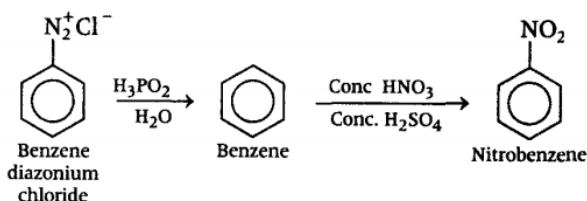


OR

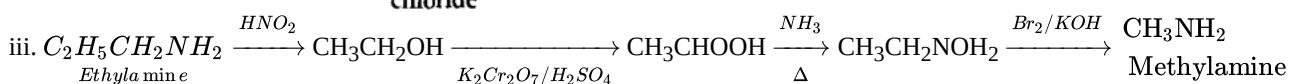
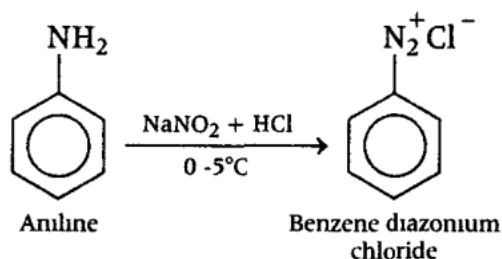


Both the compounds i.e., $FeSO_4 \cdot (NH_4)_2SO_4 \cdot 6H_2O$ and $[Cu(NH_3)_4]SO_4 \cdot 5H_2O$ fall under the category of addition compounds with only one major difference i.e., the former is an example of a double salt, while the latter is a coordination compound. A double salt is an addition compound that is stable in the solid state but that which breaks up into its constituent ions in the dissolved state. These compounds exhibit individual properties of their constituents. For e.g. $FeSO_4 \cdot (NH_4)_2SO_4 \cdot 6H_2O$ breaks into Fe^{2+} , NH_4^+ and SO_4^{2-} ions. Hence, it gives a positive test for Fe^{2+} ions. A coordination compound is an addition compound which retains its identity in the solid as well as in the dissolved state. However, the individual properties of the constituents are lost. This happens because $[Cu(NH_3)_4]SO_4 \cdot 5H_2O$ does not show the test for Cu^{2+} . The ions present in the solution of $[Cu(NH_3)_4]SO_4 \cdot 5H_2O$ are $[Cu(NH_3)_4]^{2+}$ and SO_4^{2-} .

33. i.

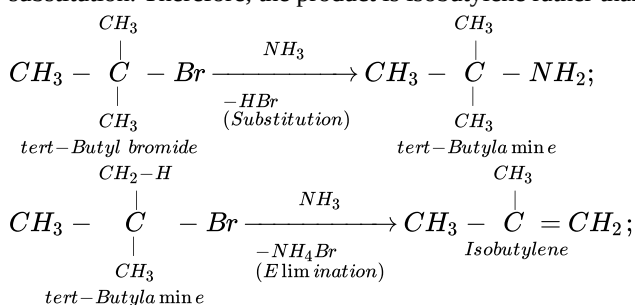


ii.



OR

i. Tert.-Butyl bromide being a 3° alkyl halide on treatment with a base (i.e., NH_3) prefers to undergo elimination rather than substitution. Therefore, the product is isobutylene rather than tert-butylamine.



ii. 1° amines containing tert-alkyl groups can be prepared by action of suitable Grignard reagents and o-methylhydroxylamine. For example,

